



Cambridge IGCSE™

CO-ORDINATED SCIENCES (DOUBLE AWARD)**0654/42**

Paper 4 Extended Theory

March 2022

MARK SCHEME

Maximum Mark: 120

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the March 2022 series for most Cambridge IGCSE™, Cambridge International A and AS Level components and some Cambridge O Level components.

This document consists of **14** printed pages.

PUBLISHED**Generic Marking Principles**

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5:

Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).

GENERIC MARKING PRINCIPLE 6:

Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

- 1 Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
- 2 The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
- 3 Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
- 4 The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
- 5 'List rule' guidance
For questions that require *n* responses (e.g. State **two** reasons ...):
 - The response should be read as continuous prose, even when numbered answer spaces are provided.
 - Any response marked *ignore* in the mark scheme should not count towards *n*.
 - Incorrect responses should not be awarded credit but will still count towards *n*.
 - Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response.
 - Non-contradictory responses after the first *n* responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Examples of how to apply the list rule

State **three** reasons.... [3]

A	1. Correct	✓	2
	2. Correct	✓	
	3. Wrong	✗	

B (4 responses)	1. Correct, Correct	✓, ✓	3
	2. Correct	✓	
	3. Wrong	ignore	

C (4 responses)	1. Correct	✓	2
	2. Correct, Wrong	✓, ✗	
	3. Correct	ignore	

D (4 responses)	1. Correct	✓	2
	2. Correct, CON (of 2.)	✗, (discount 2)	
	3. Correct	✓	

E (4 responses)	1. Correct	✓	3
	2. Correct	✓	
	3. Correct, Wrong	✓	

F (4 responses)	1. Correct	✓	2
	2. Correct	✓	
	3. Correct CON (of 3.)	✗ (discount 3)	

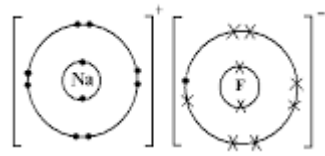
G (5 responses)	1. Correct	✓	3
	2. Correct	✓	
	3. Correct Correct CON (of 4.)	✓ ignore ignore	

H (4 responses)	1. Correct	✓	2
	2. Correct	✗	
	3. CON (of 2.) Correct	(discount 2) ✓	

I (4 responses)	1. Correct	✓	2
	2. Correct	✗	
	3. Correct CON (of 2.)	✓ (discount 2)	

Question	Answer	Marks
1(a)	sensitivity is the ability to <u>detect</u> ; and <u>respond</u> to changes in the environment ;	2
1(b)(i)	X drawn on blind spot ;	1
1(b)(ii)	H ; B ; C ;	3
1(b)(iii)	A becomes thin(ner) / AW ; E tighten / stretch / AW ; F relax ;	3
1(c)	brain ; spinal cord ;	2

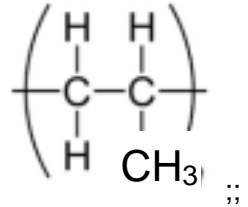
Question	Answer	Marks												
2(a)(i)	atoms with the same, proton or atomic number / number of protons ; (but) different, nucleon / mass number / number of neutrons ;	2												
2(a)(ii)	<table border="1"> <thead> <tr> <th></th> <th>potassium atom, K</th> <th>potassium ion, K⁺</th> </tr> </thead> <tbody> <tr> <td>number of protons</td> <td>19</td> <td>19</td> </tr> <tr> <td>number of electrons</td> <td>19</td> <td>18</td> </tr> <tr> <td>number of neutrons</td> <td>20</td> <td>20</td> </tr> </tbody> </table> ...		potassium atom, K	potassium ion, K ⁺	number of protons	19	19	number of electrons	19	18	number of neutrons	20	20	3
	potassium atom, K	potassium ion, K ⁺												
number of protons	19	19												
number of electrons	19	18												
number of neutrons	20	20												
2(b)	$2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$;;	2												

Question	Answer	Marks
2(c)	 ;;	2
2(d)(i)	sodium ;	1
2(d)(ii)	chloride ;	1

Question	Answer	Marks
3(a)(i)	$180\,000 / 3600 (= 50 \text{ m/s}) ;$	1
3(a)(ii)	$(KE =) \frac{1}{2} mv^2$ or $\frac{1}{2} \times 680\,000 \times 50^2 ;$ $(KE =) 850\,000\,000 \text{ (J)} ;$	2
3(b)(i)	$(v =) f \lambda$ or $250 \times 1.32 ;$ $(v =) 330 \text{ (m/s)} ;$	2
3(b)(ii)	vibrations / oscillations, of (air) particles ; rarefaction and compressions ;	2
3(c)(i)	$(V =) m / \rho$ (in any form) or $324 / 8100 ;$ $(V =) 0.04 \text{ (m}^3) ;$	2
3(c)(ii)	when the temperature of the tracks increases, the tracks will expand ; the gaps prevent buckling of the tracks / owtte ;	2

Question	Answer	Marks
4(a)(i)	<i>any three from:</i> enzymes were denatured (due to boiling) ; change in shape of active site ; enzyme is no longer complementary to substrate ; bile alone does not digest fats / fats weren't broken down into fatty acids and glycerol ;	3
4(a)(ii)	182 (seconds) ;	1
4(a)(iii)	<i>any two from:</i> bile, emulsifies the fat ; increases surface area of fat ; so there is faster, breakdown / digestion (by lipase) ;	2
4(b)	liver ;	1
4(c)(i)	large surface area ;	1
4(c)(ii)	small intestine ;	1

Question	Answer	Marks
5(a)	C_3H_6 ;	1
5(b)	alkenes ;	1
5(c)	contains a carbon to carbon double bond ;	1
5(d)	addition (reaction) / reduction / hydrogenation ;	1
5(e)	propene (aqueous bromine) changes from orange to colourless / decolourised / loses its colour / owtte propane (aqueous bromine) stays orange / no change	2

Question	Answer	Marks
5(f)(i)		2
5(f)(ii)	<p>any two from: <i>condensation</i> – one larger and one smaller molecule formed / <i>addition</i> – one large molecule formed ; in condensation polymerisation water is (often) produced (as the smaller molecule) ; AVP ;</p>	2

Question	Answer	Marks
6(a)	<p>(Wd =) mgh or $15 \times 10 \times 1.8$; (Wd =) 270 ; Joules / J ;</p>	3
6(b)	<p>acceleration ; non-constant acceleration / high then low acceleration ;</p>	2
6(c)	<p>transfer of electrons ; from the child / to the slide ; due to friction ;</p>	3
6(d)	<p>less more less ;</p>	1

Question	Answer					Marks
7(a)	organism	producer	herbivore	carnivore	quaternary consumer	3
	arctic cod			✓		
	krill		✓			
	orca			✓	✓	
	phytoplankton	✓				
;;;						
7(b)	<i>any two from:</i> by photosynthesis ; using light (energy) ; using carbon dioxide and water ;					2
7(c)	<i>any three from:</i> heat / respiration ; excretion / faeces / urea ; (named) movement / (named) metabolic processes ; not all the organism is eaten / some parts are undigested ; not all the organisms from the trophic level are eaten ;					3
7(d)(i)	selection ; environment ; alleles ; generations ;					4
7(d)(ii)	<i>any two from:</i> good blood supply ; thin (walls/surface) / short diffusion distance ; large surface area ; good ventilation ; AVP ;					2

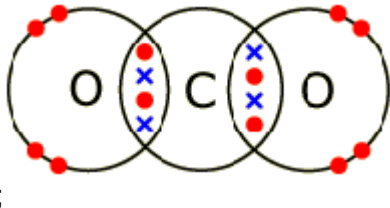
Question	Answer	Marks
8(a)	Fe / CuSO₄ pink/brown or pink or brown ; Mg / CuSO₄ metal – silver (coloured) ; solution – colourless ;	3
8(b)	$\text{Zn(s)} + \text{FeSO}_4(\text{aq}) \rightarrow \text{Fe(s)} + \text{ZnSO}_4(\text{aq}) ;;$	2
8(c)	$\text{Mg} - 2\text{e}^- \rightarrow \text{Mg}^{2+} / \text{Mg} \rightarrow \text{Mg}^{2+} + 2\text{e}^- ;;$	2
8(d)(i)	relative molecular mass of $\text{MgCl}_2 = 95$; $((95 \times 0.48) \div 24) = 1.9 \text{ (g)}$;	2
8(d)(ii)	test - lighted splint / ignite gas ; observation - (squeaky) pop ;	2

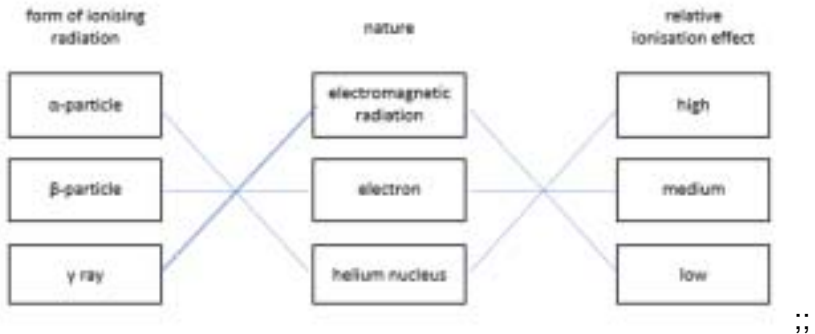
Question	Answer	Marks
9(a)(i)	extension = $3.4 - 2.2 = 1.2 \text{ (cm)}$; (F =) kx or 0.50×1.2 ; (F =) 0.6 N ;	3
9(a)(ii)	B and smallest extension ; extension is inversely proportional to spring constant ;	2

Question	Answer	Marks
9(b)(i)	<p>pd across spring = $9 - 7.5 = 1.5$ (V) ; (R =) V / I or $1.5 / 0.75$; (R =) 2 (Ω) ;</p> <p>OR</p> <p>combined resistance = V / I or $9 / 0.75$ or 12Ω ; resistance of spring = combined resistance – 10Ω ; resistance of spring = 2 (Ω) ;</p>	3
9(b)(ii)	<p>correct shape ; correct direction ;</p>	2

Question	Answer	Marks
10(a)(i)	plasmolysis ;	1
10(a)(ii)	<p><i>any three from:</i> higher water potential inside the cells than outside the cells ; water moves from high water potential to low water potential / water moves down a water potential gradient ; across a partially permeable membrane / water moves out of the cells ; ref to osmosis ;</p>	3
10(b)	<p>sucrose ; amino acids ;</p>	2

Question	Answer	Marks
11(a)	hard ;	1
11(b)	lubricant / pencils / electrodes ;	1

Question	Answer	Marks
11(c)	(graphite contains) delocalised electrons / electrons can move through the structure ; movement of charge constitutes a current ;	2
11(d)	covalent ;	1
11(e)		3

Question	Answer	Marks
12(a)	microwaves infrared ultraviolet ;;	2
12(b)	(n =) $\sin i / \sin r$ or $\sin 30 / \sin 15$; (n =) 1.93 ;	2
12(c)(i)		2

Question	Answer	Marks
12(c)(ii)	${}_{82}^{210}\text{Pb} \rightarrow {}_{83}^{210}\text{Bi} + {}_{-1}^0\beta$;	2
12(c)(iii)	identification of 190 counts per min / correct working on graph ; 22 (years) ;	2